

The reactivity series			
You should be able to:	Lesson covered	Check your understanding	
		Before	After
Metals react with oxygen to produce metal oxides. The reactions are oxidation reactions because the metals gain oxygen.			
You should be able to explain reduction and oxidation in terms of loss or gain of oxygen.			
When metals react with other substances the metal atoms form positive ions. The reactivity of a metal is related to its ability to form positive ions.			
Metals can be arranged in order of their reactivity in a reactivity series. They are put in order of their reactivity from their reactions with: <ul style="list-style-type: none"> • water • dilute acids 			
You should be able to recall the order of: potassium, sodium, lithium, calcium, magnesium, zinc, iron and copper			
Be able to recall and describe the reaction of potassium, sodium, lithium, calcium, magnesium, zinc, iron and copper with water			
Be able to recall and describe the reaction of potassium, sodium, lithium, calcium, magnesium, zinc, iron and copper with dilute acid			
The non-metals hydrogen and carbon are often included in the reactivity series.			
A more reactive metal can displace a less reactive metal from a compound.			

Keywords


- reduction
- displacement
- ions
- oxidation
- corrosion

Links between topics

- Atomic structure - know what an ion is
- Extraction of metals - using carbon in reduction reactions to extract metals
- The periodic table - group 1 and transition metals
- Bonding - know how an ionic compound forms and how formation ions relates to reactivity

Help

These websites will provide you with extra help when we are studying the topic and are useful for revision.

BBC Bitesize The reactivity series	www.bbc.co.uk/education/guides/zqjsgk7/revision	
Fuse School The reactivity series	www.youtube.com/watch?v=A-gh-feOsgo	
YouTube The reactivity series (song)	www.youtube.com/watch?v=DLlykUHHAcQ	