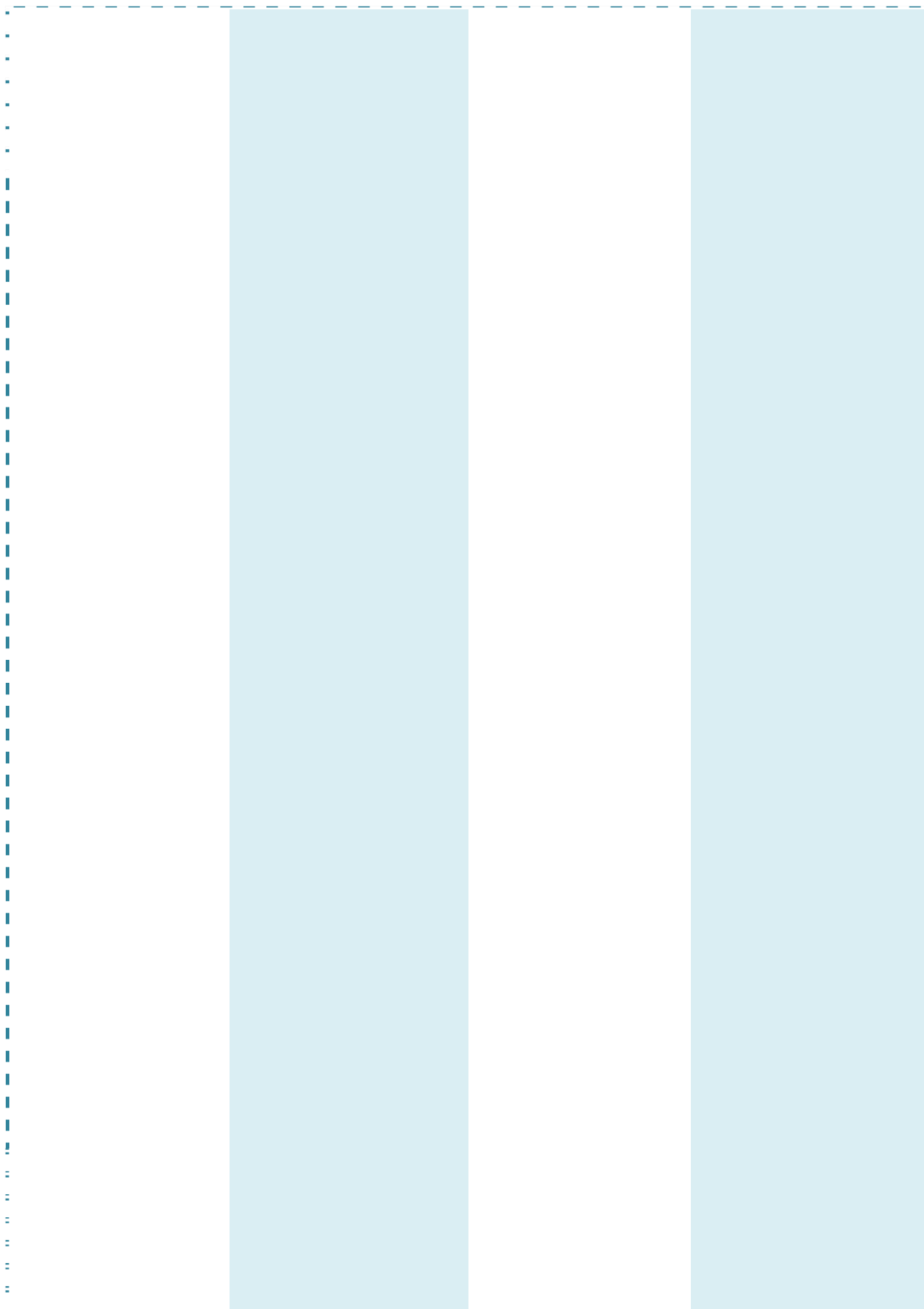


Law of conservation of mass	No atoms are made or lost in a reaction so the mass of the products is always equal to the mass of the reactants.	Relative formula mass	The sum of the relative atomic masses of the atoms in the formula.
Atomic mass	The total number of protons and neutrons in an atom.	M_r	The symbol representing relative formula mass.
Balanced equation	An equation where the number of atoms of each element is the same on each side.	Mole	A unit to measure amounts of chemicals e.g. atoms, molecules, ions and electrons.
A_r	The symbol representing relative atomic mass.	Avogadro constant	6×10^{23}
Reactants	The substances used up in a reaction.	Percentage yield	The amount of product actually made in a reaction compared with the amount that could theoretically be made.
Yield	The amount of product made in a reaction.	Concentration of a solution	Mass of solute in grams per dm^3
Atom economy	A measure of the amount of starting materials that end up as useful products.	mol/dm^3	Unit of concentration
24 dm^3	The volume of one mole of a gas at room temperature and pressure.	Limiting reactant	A reactant that is completely used up in a reaction.



Teaching notes

Before the lesson print out and cut up the cards - one set per group. The activity can easily be differentiated by removing some pairs of cards.

Students then place the cards face down on the table. (To make it easier they could group all the blue cards together and all the white cards together so that they take a card from each group to make a pair.) They take it in turns to turn over 2 cards. If they match they can keep them and have another turn. The winning team is the one with the most pairs.