

### Task 1

Use the reactivity series to predict the following reactions.

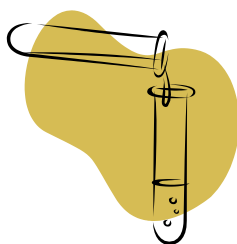
1. copper oxide + magnesium  $\longrightarrow$
2. lead nitrate + tin  $\longrightarrow$
3. lead chloride + iron  $\longrightarrow$
4. copper sulphate + calcium  $\longrightarrow$
5. aluminium + silver chloride  $\longrightarrow$
6. iron oxide + carbon  $\longrightarrow$

### Task 2

Look at this list of reactions. Use your reactivity series to predict which reactions will take place. Complete the reactions that will take place, put a cross by the ones that will not.

1. aluminium + zinc oxide
2. sodium chloride + aluminium
3. zinc sulphate + copper
4. calcium chloride + iron
5. magnesium + silver nitrate

Remember: a metal element higher in the table will displace (pinch) the non-metal element(s) (e. g. oxygen) or their salts (e.g. sulphate), from a metal lower in the reactivity series. These are displacement reactions.



## Answers

### Task 1

- 1 magnesium oxide + copper
- 2 tin nitrate + lead
- 3 iron chloride + lead
- 4 calcium sulphate + copper
- 5 aluminium chloride + silver
- 6 iron + carbon dioxide

### Task 2

- |    |                             |   |                            |
|----|-----------------------------|---|----------------------------|
| 1. | aluminium + zinc oxide      |   | zinc + aluminium oxide     |
| 2. | sodium chloride + aluminium | x |                            |
| 3. | zinc sulphate + copper      | x |                            |
| 4. | calcium chloride + iron     | x |                            |
| 5. | magnesium + silver nitrate  |   | silver nitrate + magnesium |